Chemistry I   Chapter 22 Quiz   Name_____

<u>SHOW ALL WORK</u>

1.   At 25°C, 7.8×10^{-5} mole of silver chromate dissolves in 1.0
     liter of water. What is the K_{sp} of Ag_2CrO_4?

2.   The K_{sp} of CaF_2 is 3.9×10^{-11} at 25°C. What is the
     concentration of Ca^{2+} and F^- in the saturated solution?

3.   Will a precipitate form if 10.0 mL of 0.010M $AgNO_3$ and 10.0
     mL of 0.00010M NaCl are mixed? Assume the final volume of
     the solution is 20.0 mL. For AgCl, $K_{sp} = 1.7 \times 10^{-10}$.