Chemistry I   Chapter 22 Quiz #2   Name_____

<u>SHOW ALL WORK</u>

1.   At 25°C, 3.6×10^{-19} mole of nickel (II) phosphate dissolves
      in 2.0 liter of water. What is the K_{sp} of $Ni_3(PO_4)_2$?

2.   The K_{sp} of $Pb(IO_3)_2$ is 2.6×10^{-13} at 25°C. What is the
      concentration of Pb^{2+} and IO_3^- in the saturated solution?

3.   Decide whether a precipitate will form when 40.0 mL of
      0.0010 M K_2CO_3 is added to 60.0 mL of 1.0×10^{-6} M $AgNO_3$.